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Stoichiometry can be used to determine the chemical formula of a compound by studying the relative amounts of elements present or can be used to study the relative amounts of compounds that are consumed and produced during a chemical reaction.

Chapter 3 Stoichiometry

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Abstract and Figures This case study presents a hypothetical situation that utilizes the principles of stoichiometry and limiting reagents to a chemical production problem. It is a collaborative...

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Stoichiometry is one of the most important subjects in general chemistry. It is typically introduced after discussing parts of the atom and unit conversions. While it's not difficult, many students get put off by the complicated-sounding word. For this reason, it may be introduced as "Mass Relations."

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Stoichiometry The atomic ratios in each compound are also the relative number of atomic mass units of its elements. The first example is nitrous oxide (N₂O), as shown in Table 1. The relative masses were obtained by multiplying the atomic ratios and atomic masses.

Stoichiometry

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in the Study Guide for Content Mastery Study Guide for Content Mastery - Student Edition Stoichiometry is the tool for answering these questions. Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemi-cal reaction is called stoichiometry.

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