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Chemistry Of Copper Pre Lab

Chemistry of Copper
Lab 3 Pages 109 - 115
Pre-lab pages 111 -
112 Post lab questions
page 114 - 115 .

Introduction • Copper
is found in group 11,
MW = 63.456 • Shiny

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(orange/red color),

Malleable, Ductile •

Oxidizes in air (turns a
green color - patina)

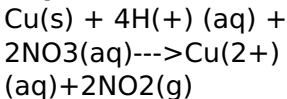
Chemistry of Copper

After sulfuric acid was added to the beaker, copper was found as copper ions with a 2+ charge instead of the previous copper(ii) oxide form. 6. In the final step of the lab when the copper precipitate was washed, zinc ions were

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removed. The previous reaction that took place involved aqueous copper(ii) sulfate and solid zinc.

Copper Lab - AP Chemistry Lab Reports



+2H₂O(l) Acid base:
NO₃ because after it combines with H it becomes an acid.

Copper forms many

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different compounds in
this experiment.

Identify the oxidizing
agent in the conversion
of copper metal to
copper(II) ion.

chemistry of copper lab 28 Flashcards | Quizlet

Chemistry Of Copper
Pre Lab • Copper(II)
sulfate is a desiccant. •
Copper sulfate is a
commonly included
chemical in children's.

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Page 7/24

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chemistry sets and is often used in high school crystal growing and copper plating experiments. • A very dilute solution of Copper sulfate is used to treat aquarium fish of various parasitic infections.

Chemistry Of Copper Pre Lab Answers

$\text{Cu}(\text{OH})_2$ - A black solution was formed.

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After being heated, the solution turned into water and a black precipitate. All of the black precipitate would collect at the bottom because it is denser than water. $\text{CuO} + \text{H}_2\text{SO}_4$ - The solution turned from a black solution into a greenish blue solution.

Copper Lab - AP Chemistry Labs

1. Copper
2. Add HNO_3
3. Warm solution to

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dissolve copper further

4. Place DI water and dissolve solution while stirring slowly 5. Add NaOH and stir until precipitation is complete (test pH to make its basic) 6.

Boiling chips + heat basic solution--> precipitate at bottom and colorless liquid on top 7.

Experiment 3-Copper Chemistry Pre and Post Lab

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(s) is heated, Copper (II) oxide and water are formed. Write a

balanced equation for the reaction. $\text{Cu}(\text{OH})_2(\text{s}) \rightarrow \text{CuO}(\text{s}) + \text{H}_2\text{O}(\text{g})$

10. When sulfuric acid and copper (II) oxide are allowed to react, copper (II) sulfate and water are formed.

Write a balanced equation for this reaction. $\text{H}_2\text{SO}_4(\text{aq}) + \text{CuO}(\text{s}) \rightarrow \text{CuSO}_4$

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Chemical Reactions of Copper and Percent Yield

Types of Reactions:
The Copper Cycle In
this laboratory
experiment, students
will perform a series of
reactions known as the
copper cycle. The
reaction series includes
single replacement,
double replacement,
synthesis, and
decomposition
reactions.

Download Free Chemistry Of

Copper Pre Lab **Types of Reactions: The Copper cycle**

In this experiment, you will perform and observe several reactions of copper.

This is a cycle of reactions, because you start and end with the same substance, copper metal. In the first reaction, copper metal is oxidized by nitric acid to form copper (II) nitrate, $\text{Cu}(\text{NO}_3)_2$

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Copper Pre-Lab **Experiment 11 - A Cycle of Copper Reactions**

The objective of this lab was to fully carry out five reactions of copper, and to observe and understand the methods behind each reaction. The copper first underwent a redox reaction with nitric acid, resulting in a light blue fluid. Then, NaOH was added in order to create a copper precipitate, creating a

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darker blue cooler.

**Chemical Reactions
of Copper Lab by
Natalie Dickman on**

...

Lab #6 Chemical
Transformations of
Copper Introduction:
Copper was one of the
first metals to be
isolated, due to the
ease of separating it
from its ... copper solid
on a pre-weighed
glassine weighing
paper, place the paper

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with its precipitate on a watch glass and place in a drying oven for 5 - 10 minutes. ...

Lab #6 Chemical Transformations of Copper

The Copper Lab demonstrates stoichiometry in chemistry.

Stoichiometry is helpful in calculating the amount of an element or compound in chemical reactions. For

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the lab, stoichiometry was used to predict the amount of copper that would be left over.

Copper Lab - AP Chemistry

Experiment 3

Prelaboratory

Assignment: Chemistry of Copper. 1. 1. Review net-ionic equations 3.2, 3.4, 3.7, 3.9, and 3.11.

-a. Three of the equations represent oxidation-reduction reactions. Identify the

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three equations and indicate the oxidizing agent in each. ... If 0.0169g of copper is the recovered after the series of reactions in this ...

Print Experiment 3 Prelaboratory

Assignment: Chemistry of ...

In step 1 of the experimental procedure, copper metal is added to concentrated nitric

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acid. The reaction between copper metal and concentrated nitric acid is an oxidation-reduction reaction that is somewhat complicated. •4 HNO₃. (aq)+ Cu(s)→ Cu(NO.

Chemistry 2A Lab Manual

Chemistry of Copper
Lab 3 Pages 109 - 115
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Introduction • Copper

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(orange/red color),
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Oxidizes in air (turns a
green color - patina)

Chemical Reactions of
Copper and Percent
Yield

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Copper was put
through a series of
several reactions, until
solid copper was the

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last remaining product.
The initial mass of copper was recorded, and then compared with the mass of recovered copper at the end of the experiment.

Chemistry Lab Report (Copper Cycle) - Sarah Jackson

The original copper sample contained a specific number of moles of copper, and

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that amount of copper was present in every reaction, precipitate, and solution. Thus, the same number of moles of copper was produced in the final reaction because the amount of copper remained constant throughout the experiment.

Lab Report on copper cycle - LinkedIn SlideShare

Copper sulfate

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electrolyte solution
(200 g $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$
+ 25.0 mL
concentrated H_2SO_4
solution in enough
distilled or deionized
water to make 1.0 L of
solution) Coin cleaning
solution (1 tsp table
salt dissolved in $\frac{1}{4}$ cup
vinegar) Pre-1982
penny (contains 95%
copper, versus only
2.5% copper
post-1982) Dime (any
year)

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